



PimaCommunityCollege
West Campus

CHM 151IN General Chemistry I Syllabus for Spring 2009

Course Information:

Course Prefix/Number: **CHM 151 IN**

Semester: **Spring 2009**

Class Days/Times: **MW 10:20-1:30**

Credit Hours: **5.0**

Course Title: **General Chemistry I**

CRN (Section Code): **25086**

Site/Room: **SCI K-221**

Teaching Format: **Integrated Lecture and Lab**

Instructor Information:

Name: **David A. Katz**

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(Select Pima Chem Courses, then CHM 151 or General Chem Survival Manual)

Availability: **Office hours: MW 9:00-10:00 a.m.; 2:00-3:00 p.m. TTh 11:00-12:00 a.m.; 3:30-4:30 p.m.**

Generally, in addition to my office hours, I am in the office at least 30 minutes before or after class (if I am not in the lab). I am also available by appointment.

Instructional Materials:

Required Text: Kotz, John C., Treichel, Paul M., and Townsend, John R., **Chemistry & Chemical Reactivity**, 7th Ed., Thomson-Brooks/Cole, 2009.

Availability: Pima C.C. bookstore and other academic bookstores in town. (The PCC Bookstore can be accessed and books ordered via the Internet at www.Pima.bkstr.com.)

e-book at www.ichapters.com, search for ISBN 0495387037 (The e-book is approx ½ price)

Online at www.amazon.com or www.half.com (Both new and used books available)

Laboratory Manual: Selegue, Thomas, **General Chemistry in Action**, Hayden McNeil, 2008

Important: You must have access to a computer with Internet connections. Important materials for both reading and problem assignments for class are available at www.chymist.com under the submenu "General Chemistry Survival Manual". This syllabus, along with some additional readings and references, is also available under the submenus "Pima Chem Courses" and "Chem 151".

Required: A **scientific calculator**. Cell phones and PDA's are not permitted in place of calculators in class.

Optional: Banks, Alton J., **Student Solutions Manual for Chemistry and Chemical Reactivity**, 7th Ed.

Course Description:

THIS COURSE HAS BOTH A MATH AND A CHEMISTRY PREREQUISITE:

Minimum Math prerequisite: MAT 122 with a grade of C or better.

Chemistry prerequisite: completion of CHM 080 or CHM 130 with grade of C or better, or placement into CHM 151 on the Chemistry assessment test.

Course Corequisite: CHM 151LB

CHEMISTRY 151 is an introduction to the foundations of chemistry for upper-level sciences and engineering. The content of this course includes both principles and problems in the areas of atomic structure, introductory nuclear chemistry, chemical bonding, reaction stoichiometry, behavior of gases, reactions in solutions, solution concentrations, and an introduction to thermochemistry.

Problem solving is a necessary part of the course, students should review basic mathematics and algebra with emphasis on exponents and roots, manipulations of algebraic equations, and construction of graphs and tables.

The use of scientific calculators is permitted for all aspects of this course.

Course Objectives:

Upon completion of the course, the student will be able to do the following:

1. Apply the concepts associated with measurement, units, significant figures, dimensional analysis, etc. to chemistry related calculations.
2. Describe the scientific method in principle and practice.
3. Classify matter and its changes.
4. Demonstrate familiarity with the historical development of the atom, its component parts, atomic mass and isotopes.
5. Associate the wave-particle duality of the electromagnetic spectrum with the wave-particle view of the electron.
6. Describe the pertinent aspects of the quantum-mechanical view of the atom.
7. Identify the electron configuration of atoms and ions.
8. Demonstrate the ability to use the Periodic Table to extract a range of information including number of valence electrons, orbital information, metal/nonmetal characteristics, periodic trends and the reason for similarity in chemical characteristics.
9. Describe the concepts associated with ionic and covalent bonding. Predict the nature and formula of bonded atoms.
10. Draw Lewis Structures and determine the molecular shape for given compounds.
11. Determine the polarity of a given molecule.
12. Apply Valence Bond Theory to determine the hybridization of bonded atoms.
13. Balance and extract information from chemical equations.
14. Identify simple reaction types.
15. Predict solubility and write supporting equations.
16. Perform solution concentration calculations.
17. Perform stoichiometry calculations for a range of systems including limiting reagent and aqueous solutions.
18. Describe the thermodynamic variable of enthalpy.
19. Calculate reaction enthalpy's from Heats of Formation data.
20. Calculate the energy associated with temperature changes including the concept of calorimetry.
21. Describe and apply Kinetic-Molecular Theory
22. Describe the various intermolecular forces and identify their existence/function in a variety of systems.
23. Apply the Ideal Gas Law to systems of gases including the concept of gas mixtures.
24. Distinguish between ideal and non-ideal systems.
25. Describe the various properties associated with liquids, e.g. vapor pressure, surface tension, viscosity, etc.
26. Describe the concept of a solution.
27. Perform concentration calculations.
28. Calculate/describe the solution activity associated with colligative properties. (Optional)

Lecture Outline

The following outline presents topics and subtopics in the order they will be discussed in class along with reading assignments and problem assignments. The course content is integrated by topic and does not necessarily follow the book in a chapter by chapter and page by page order.

Read chapters before the material is discussed in lecture.

Problem assignments list the problem sections at the end of each chapter. Applicable problems should be attempted and solved after methods of problem solving are demonstrated and explained in class. Answers to odd number problems are located in the back of the textbook. An optional student solutions manual is available.

Lecture Topics	Reading Assignment	Problem Assignment
<p>Pre-Class Assignments Math Review (You should be able to solve these problems to succeed in this course.) Significant Figures, Exponents, and Scientific Notation</p>	<p>Math Review (on web site)</p> <p>Significant Figures, Exponents, and Scientific Notation (on web site)</p>	<p>Math Review All problems</p> <p>Significant Figures, Exponents, and Scientific Notation All problems</p>
<p>1. Introduction and Review</p> <ol style="list-style-type: none"> Chemistry, life, and the universe Scientific method Matter and its properties Density Elements, compounds, and mixtures Element symbols Names and formulas of inorganic compounds How big is an atom? : Measurement and the SI system How hot is hot?: Temperature Precision, accuracy, experimental error, and standard deviation Dimensional Analysis problem solving <p>Note: This information will be treated as a review in class as it is part of the prerequisites for this course</p>	<p>Chapter 1 pages 1-43</p> <p>Metric System (on web site)</p> <p>Temperature (on web site)</p> <p>Element Symbols (on web site)</p> <p>Formula Writing and Nomenclature (on web site)</p> <p>Problem Solving by Dimensional Analysis (on web site)</p>	<p>Chapter 1 pages 20-23 1-11, 13-37 (odd nos.)</p> <p>Pages, 43-49 1-59 (odd nos.)</p> <p>Temperature All problems</p> <p>Formula Writing and Nomenclature All problems</p> <p>Problem Solving by Dimensional Analysis All problems</p>
<p>2. Atoms, Molecules and Intro to Nuclear Chemistry</p> <ol style="list-style-type: none"> Identifying elements : A brief history of Chemistry to 1800 John Dalton: The atomic theory and atomic mass Discovery and characterization of the electron The proton Rutherford and the nuclear atom The neutron Isotopes Atomic weights Atomic numbers History of the periodic classification The periodic law The modern periodic table Radioactivity Radioactive decay Nuclear stability Half-life: Rates of radioactive decay Radioisotope dating Nuclear transformations Extending the periodic table Biological effects of radiation Applications of radioisotopes 	<p>An Illustrated History of Alchemy and Chemistry (on web site: History of Chem)</p> <p>Chapter 2 Pages 50-82</p> <p>Milestones in the Development of Chemistry and the Modern View of Atoms and Molecules Pages 338-347</p> <p>Chapter 23 Pages 1060-1090 Pages 338-347</p> <p>Prospects for Further Considerable Extension of the Periodic Table (on web site)</p> <p>Web Elements Periodic Table http://www.webelements.com</p>	<p>Chapter 2 Pages 100-111 1-51 (odd nos.), 85, 87, 91, 115, 129</p> <p>Chapter 23 pages 1090-1095 7, 9-19 (odd nos.), 29-43 (odd nos.), 53</p> <p>Chapter 10 pages 488-495 1-7 (odd nos.), 9 (draw isomers only), 15, 25, 31-35 (odd nos.), 39, 41, 51</p>

Lecture Topics	Reading Assignment	Problem Assignment
2. Atoms, Molecules and Intro to Nuclear Chemistry (continued) v) Intro to organic compounds w) Intro to functional group compounds	Chapter 10 Pages 442-478	
3. Chemical Formulas, Moles and Chemical Reactions a) Formula mass/Molar mass b) The Mole and Avogadro's number c) Percent composition d) Empirical and molecular formulas e) Hydrates f) The meaning of a chemical equation g) Balancing equations h) Intro to Chemical Equilibrium i) Types of Reactions j) Acids and bases k) Electrolytes l) Net ionic equations m) Intro to redox reactions n) The activity series	Chapter 2 pages 82-100 Chemical Formulas and Formula Mass Calculations (on web site) Writing Chemical Equations (on web site) Chapter 3 pages 112-152	Chapter 2 pages 100-111 53-83 (odd nos.), 89, 93, 95, 101-113 (odd nos.), 117, 119, 121, 127, 133, 141, 153 Chemical Formulas and Formula Mass Calculations All problems Writing Chemical Equations All problems Chapter 3 pages 152-157 1-69 (odd nos.)
4. Stoichiometry a) Mass relationships in reactions b) Limiting reagents c) Theoretical and percent yields d) Reactions in solution e) Water as a solvent f) Electrolytes g) Solution terminology h) Percent i) Molarity j) Dilution k) pH l) Solution stoichiometry	Chapter 4 Pages 158-195	Chapter 4 Page 195-207 1-73 (odd nos.), 77-113 (odd nos.)
5. Energy and Chemical reactions a) Heat and energy b) Specific heat c) Energy and changes of state d) The first law of thermodynamics e) Energy changes for chemical reactions f) Calorimetry g) Hess's Law h) Standard enthalpy of formation i) Energy change for a reaction	Chapter 5 Pages 208-242 The Chemistry of Fuels and Energy Resources Pages 254-266	Chapter 5 Pages 242-253 3, 6, 7-57 (odd nos.), 67- 81 (odd nos.), 85, 89 The Chemistry of Fuels and Energy Resources Pages 266-267 1, 5, 7, 9, 11, 13, 17

Lecture Topics	Reading Assignment	Problem Assignment
<p>6. Electronic Structure of Atoms and the Periodic Table</p> <ul style="list-style-type: none"> a) Electromagnetic radiation b) Atomic spectra c) The quantum theory d) The photoelectric effect e) The Bohr model f) Electron waves g) The uncertainty principle h) The quantum mechanical atom i) Quantum numbers j) Energy levels and orbitals k) Electron distributions l) Magnetic susceptibility m) Hund's rule n) The periodicity of electron configurations o) Electron configurations and ions p) Atomic Size q) Ionization energy r) Electron affinity s) Ionic radii t) Atomic structure and reactivity 	<p>Chapter 6 pages 268-296</p> <p>Chapter 7 pages 304-331</p>	<p>Chapter 6 pages 297-302 1-17 (odd nos.), 23-35 (odd nos.), 53, 59, 61, 65, 67, 69, 71, 77</p> <p>Chapter 7 pages 332-337 1-9 (odd nos.), 17-33 (odd nos.), 36-43, 47, 49, 51, 55, 61, 63, 69</p>
<p>7. Chemical Bonding and Molecular Geometry</p> <ul style="list-style-type: none"> a) Lewis dot symbols of elements b) The ionic bond c) The covalent bond d) Multiple bonds e) The octet rule f) Bond distances and bond strength g) Bond polarity h) Electronegativity i) The metallic bond j) Lewis structures k) Bond energies from Lewis structures l) Molecular shapes m) VSEPR Theory n) Polarity of molecules o) Valence Bond Theory p) Hybrid orbitals q) Intro to Molecular orbital theory 	<p>Chapter 8 Pages 348-395</p> <p>Chapter 9 pages 404-433</p> <p>The Chemistry of Modern Materials Pages 657-663</p>	<p>Chapter 8 pages 395-403 1-73 (odd nos.), 85</p> <p>Chapter 9 pages 434-441 1-35 (odd nos.), 39, 41, 45</p>

Lecture Topics	Reading Assignment	Problem Assignment
8. Gases, Liquids, Solids and Intermolecular Forces a) Gases vs. liquids vs. solids b) Pressure c) Boyle's Law d) Charles' Law e) The combined gas law f) The Ideal Gas Law g) Density and Molar mass calculations h) Avogadro's Law: Stoichiometry in gas reactions i) Dalton's Law j) Graham's Law k) The Kinetic-Molecular Theory l) Real gases m) Intermolecular forces n) Surface tension o) Viscosity p) Evaporation q) Heat of vaporization r) Vapor pressure s) Boiling and freezing points t) Crystal structure u) X-ray diffraction v) Types of crystals w) Unit cells x) Phase changes and phase diagrams	Chapter 11 pages 514-545 Chapter 12 Pages 554-581 Chapter 13 Pages 588-610	Chapter 11 pages 546-553 1-51 (odd nos.), 57, 59, 63, 65, 67, 81 Chapter 12 pages 581-587 1-33 (odd nos.), 36, 37-47 (odd nos.) Chapter 13 Pages 610-615 3-11 (odd nos.), 19-25 (odd nos.)

Course Requirements:

The final course grade will be based on quizzes, hour exams, a final exam and laboratory. The approximate percent weight of each is given below:

Quizzes	15%
Exams	40%
Final Exam	20%
Lab	25%

CHM 151IN Course Policies and Information

Homework/Problem Assignments

Each type of problem or calculation is explained in class. Problem assignments, as listed in the lecture outline, are your responsibility to work out and review. Specific problems may be specified or deleted by your instructor. Since student solution manuals are available, problem assignments are not collected or graded. After a particular type of problem is reviewed in class, your instructor will schedule a quiz on that type of problem.

Proper problem solving means that each problem requires a clear and organized set-up. If the problem involves a formula, you must show the formula, substitution of fixed values and variables, and the proper units. Problems that do not utilize a formula must be clearly set up with all the proper units. Final answers should be underlined or circled and must include the proper units. Whether or not it is stated on exams or quizzes, you cannot get full credit for a problem without an organized set-up. The final numerical answer to any problem, on a quiz or exam, is worth one point.

Calculators

You are required to have a scientific calculator (graphing functions not required) for quizzes and exams. You may not use the calculator function on your cell phone or on a Palm Pilot or equivalent PDA.

Computer

This course is taught as a web-enhanced course. You must have access to a computer. This syllabus, along with reading material and problem sets are available on my web site at <http://www.chymist.com>. The syllabus and some additional readings are located on the *Pima Chem Courses* link under Chem 151. Reading material and problem sets are located under the heading of **General Chemistry Survival Manual**.

You may need to access additional information for this course on the Internet or in reference books.

Quizzes

There will be frequent quizzes during the semester. All quizzes are announced in class. Each quiz will cover a specific topic or assignment. Questions and problems on quizzes are aimed at showing basic skills in problem solving.

All quizzes have the same weight, even if the point count on particular quizzes differs. Every effort is made to grade and return quizzes by the next class. Grades on quizzes are calculated as fractions or percentages.

At the conclusion of each quiz, an answer key is distributed to the class and the quiz is reviewed. Once the first answer key is distributed, the quiz is considered to be over and anyone arriving late will be considered to have missed that quiz.

There is no curve applied to any quiz.

Assuming at least 6 quizzes have been given during the semester, if you have taken all the quizzes, your lowest quiz grade will be dropped when calculating your final quiz average.

If you miss more than 30% of the quizzes during the semester, then the weight of your quiz average, in calculating your final class average will be discounted by 50%. If you miss more than 50% of the quizzes, then none of your quizzes will be counted in your final class average.

Exams

There will be three exams (with a possibility of a fourth exam) during the semester. The exams cover information from the appropriate chapters in the textbook, information discussed in class, classroom demonstrations, and any supplementary readings or related material assigned. Questions and problems are designed to show understanding of course material and may be more complex than those problems which appeared on quizzes.

All problems will require proper organized set-ups and calculations. Some questions will be in multiple choice format and some questions may require written answers or short discussions (short essay type questions). Every effort is made to grade and return exams in about one week to ten days (about three class periods). Grades on exams will be calculated as percentages based on total possible point scores.

Exams are given during normal class time. You are expected to complete the exam within that time period. If you need extra time, please request taking the exam in the Testing and Assessment Center.

There is no curve applied to any exam.

Since information that appears on quizzes also appears on exams, if you have completed all the exams during the semester and your exam average is higher than your quiz average, then your quiz average will be dropped from your final grade calculation.

If you have completed all the exams during the semester, your lowest exam grade will be discounted by 50% when calculating your exam average for the semester.

Final Exam

There will be a final exam at the end of the semester. The Department of Chemistry utilizes the American Chemical Society (ACS) General Chemistry Exam as the final exam. The exam is cumulative, and will include concepts and calculations from the entire semester. Conceptual understanding of course material is important for the final exam. There are study guides for the ACS exam in the library, but do not rely on the study guide alone in preparing for the exam. In addition, there may be a sheet of supplementary questions or problems added to the exam if the class, as a whole, is not mastering specific topics.

The grades on the final exam are adjusted for both local and national norms. If your final exam grade is significantly higher than your exam and quiz average during the semester, extra weight will be given to the final exam.

The final exam utilizes a confidential test booklet. Any writing in the test booklet will reduce your final score by 10 points. If the writing in the test booklet is considered to be extensive, you will receive a grade of zero for the final exam.

The final exam is required for completion of this course. Failure to take the final exam will result in a grade of zero.

If you miss the final exam for some valid reason, you must contact your instructor within 24 hours after the exam or the assigned grade of zero will not be removed. Your instructor will require documentation of your excuse.

Laboratory Experiments and Laboratory Reports

This is an integrated class and there are laboratory experiments or activities scheduled for almost every class.

You are expected to read each experiment and check the safety precautions for all chemicals used in the experiments before coming to class. Some experiments may be on the web site. You are expected to download and read those experiments or activities before coming to class.

Laboratory reports follow the format outlined later in this syllabus. Some downloaded experiments may contain data pages which can be used in the laboratory reports. Data analysis calculations, graphs, and questions must be completed for each laboratory report. Only one report is required for each laboratory group unless otherwise specified by your instructor.

Reports are due no later than **one week** after the experiment is completed.

Laboratory reports are graded on a 10 point scale based on neatness, completion of data, and answers to questions. Missed or incomplete experiment reports may be graded as a zero.

Submitting Work and/or Contacting Your Instructor

Any take-home quizzes, exams, or projects should be submitted in writing no later than the dates specified. Depending on the type of assignment, late papers may not be accepted. For any assignment that has a flexible deadline, late papers will be downgraded by 10 points the rest of the day after class is over. An additional 10 points will be deducted until the next class period. After that, a grade of zero will be recorded for that assignment.

You may submit your assignment via the Internet if you cannot be in class on the day it is due. The time stamp on the message, when it is received, will determine the date. **You should receive a reply confirming that your emailed assignment was received within 48 hours of submitting it.**

If you contact your instructor by email, please be advised that email occasionally gets misdirected, can end up being blocked by a spam filter, or lost in cyberspace. (Blank subject lines or subjects such as "Hello" may go directly into a trash file.) **You should receive a reply confirming that your emailed message was received within 48 hours of submitting it.** It is your responsibility to make sure that the message was received.

Make-up Policy

THERE ARE NO MAKE-UP EXAMS OR QUIZZES no matter how valid your excuse may be. All exams and quizzes are scheduled with advance notice and they are given once only. **THIS INCLUDES PARTS OF A QUIZ OR EXAM MISSED AS A RESULT OF LATENESS TO CLASS.** If you miss a quiz or exam, it does not count against you (i.e., you will not get a grade of "zero"), but your grade will be calculated differently from the percent weights stated in the course requirements.

MISSING THE FINAL EXAM WILL RESULT IN A GRADE OF "ZERO" AND CAN RESULT IN YOUR FAILING THE COURSE.

If you have a previous scheduled appointment or event or know in advance that you will be late or not be present for a quiz or an exam, arrangements can be made for you to take that quiz or exam under special conditions, such as taking the exam one day earlier than the class. Such requests must be made at least three days before the scheduled test day.

If you have a conflict with the final exam, it must be resolved by the last day of class during the semester. Failure to resolve a conflict will not excuse you from the final exam.

If you desire more time to complete an exam or quiz or just want to take that test in a different-than-the-class environment, you may request to take the test in the Testing and Assessment Center in the C Building. Please make that request at least one class before the scheduled test.

Academic Integrity

Violations of scholastic ethics are considered serious offenses by Pima Community College, the Department of Chemistry and by your instructor.

Cheating on exams or quizzes will result in a grade of "zero" for that exam or quiz, and, at the instructor's discretion, possibly an F for the course. The zero will be calculated into your final grade point average for this course. Any zero grade assigned for cheating will remain as part of your average and will not be dropped or discounted.

Duplicate answers to questions or problems on quizzes and/or exams from individuals sitting next to one another can be considered as circumstantial evidence of cheating.

Occasionally, you may receive a take-home worksheet, quiz or exam. Such assignments can be completed by conferring with other students, researching information in textbooks or reference books, or on the Internet. It is expected that the final work turned in will be essentially your own, but it may be similar to that of other students you may have worked with. A large number of identical papers in the class will invalidate that assignment.

If your instructor allows you to remain in class, any cheating will be calculated into your final grade. Students who have received zeros for cheating will not be given W grades if they have not withdrawn from the course by the official withdrawal date.

Cheating may be reported to the Academic Dean, even if the student(s) involved have withdrawn from the course.

Students may consult the PCC Student Handbook sections on student code of conduct, on scholastic ethics and on the grade appeal procedure. Copies are available at PCC campus libraries and at <http://www.pima.edu/~coadmissions/studresp.htm>.

Attendance

As this is an integrated class, attendance is required for this course. Most of the important material that will be on the exams and quizzes is discussed in class. **Not everything is in the textbook.** Relevant information and applications of course material, both in PowerPoint presentations, as well as demonstrations, are also presented in class. You cannot pass the exams by reading the textbook and supplementary materials only. If you miss a class, your instructor can tell you what material was covered and summarize any discussions that took place, however, instructors do not have a set of formal lecture notes you can copy. It is your responsibility to get detailed notes from one (or preferably two) classmate(s).

The laboratory component is also a required part of this class. You are required to complete the laboratory experiments. Absence from class also counts as absence from laboratory.

If you are absent from class for an extended period due to illness, an accident, or another valid reason, please have someone contact your instructor.

Extra Credit

There is no extra credit.

You must demonstrate that you have mastered a substantial amount of the course material to obtain a passing grade.

Lateness

Lateness to class is disruptive to the other students. Please make every effort to get to class on time. If you do arrive late for a class, please come into the room quietly, find a seat, and get organized for class. Please avoid walking across the front of the classroom, disrupting the lecture or discussion. Try to walk around the back of the room.

If work or other circumstances will prevent you from getting to class on time, please see your instructor as soon as possible to discuss the problem.

Storms/Traffic or Other Problems

In the event of a severe storm, other major weather problem, a severe traffic problem, or a power outage, the area may experience transportation disruptions and traffic delays. If an exam is scheduled or a major assignment is due on a day when there is a severe weather (this excludes normal rainfalls), a major traffic problem, or other area problem, the exam or assignment deadline will be postponed until the next class.

Classroom Behavior:

Because of insurance limitations, non-registered visitors are not allowed at class sessions or on field trips.

Possession of drugs, alcohol or firearms on college property is illegal.

Eating, drinking, smoking and soliciting are not allowed in classrooms.

Pets (service animals excepted), telephones, pagers and other electronic devices that distract students are not permitted in classrooms. Please turn off these devices during classes.

Students are expected to exhibit courteous and respectful behavior in class. You are here for the purpose of furthering your education. Students who are disrespectful of others or who are creating disturbances that interfere with the conduct of the class or the learning of others will be asked to leave.

Withdrawals

Students may withdraw from class at any time during the first 2/3 of the semester without instructor permission and without incurring any grade penalty. Please be kind enough to inform your instructor if you withdraw.

Students who are not regularly attending class and who have not submitted any assignments nor taken any quizzes or exams by the census date are assumed NOT to be participating in the class may be withdrawn by the instructor.

Please be sure to withdraw yourself before the 2/3 semester deadline if you do not expect to complete the class; otherwise you may receive a grade of "F" for the course and may affect your academic standing at the college as well as any financial arrangements such as loans or scholarships.

Your instructor has the ability to enter a **late withdrawal** grade (W) at the end of the semester as a final grade for the course if you request it, in writing by the date specified in this syllabus, and have not withdrawn by the normal two-thirds mark of the semester. Once final grades have been entered, you cannot retroactively request a D or F to be changed to a W.

If you have not attended class, and have not withdrawn by the 2/3 semester date, your instructor will not honor a late withdrawal request.

Incomplete grades are only given when a significant amount of class has been missed for medical or other legitimate reasons. (See information which follows)

ADA Compliance

Pima County Community College District strives to comply with the provisions of the Americans with Disabilities Act and Section 504 of the Rehabilitation Act. Students with disabilities requiring special accommodations must notify the instructor of this need or directly contact the Disabled Student Resources Office on your campus at the beginning of the semester.

Workload

Students are expected to spend the normal amount of time required for a college course attending class sessions, doing assignments and research, reading and preparing for exams. The standard Carnegie Unit of college credit assigns 1 credit hour for each 15 hours of class time and assumes that students spend two hours working outside the classroom for each hour of classroom instruction. For a three-credit course, this translates to 135 hours per semester or an average of nine hours per week for a 15/16-week semester.

Spring 2009 Calendar of Important Dates

Jan. 20	Spring classes begin
Feb. 2	Last day to withdraw with a refund
Feb. 26-27	Rodeo Days – no classes
March 5	45 th day
March 16-22	Spring Break – no classes
April 9	Withdrawal deadline
May 12	Last day of classes – If you want a late withdrawal grade, you must request it, in writing, by today.
May 13-19	Final exam week

CHM 151IN Grading System/Policies

Your final grade will be a weighted average of your work during the semester and are calculated as follows:

A	=	100-90%
B	=	89-80%
C	=	79-70%
D	=	69-60%
F	=	below 60%

The actual percentage may vary based on a final class distribution, but will not be higher than these percentages. At most, the above stated percentages may vary by up to 5 points. I do not grade on a curve.

When calculating final grade averages, assuming you have taken all the quizzes (a total of 6 or more), your lowest quiz grade will be dropped. Also, if you have taken all the quizzes and exams, your lowest exam score will be discounted by 50% so it will have a minimum effect on your grade average.

Any grades of “zero” assigned for cheating or plagiarism will not be dropped or discounted and will be calculated into your final average. If the ethics violation is considered severe, the “zero” can override all other grades in this course.

A grade of 70 or better on the final exam will guarantee that you will receive a grade no lower than a “C” in this course. Please note that the final exam is cumulative and obtaining a grade of 70 means that you need to understand the material from the entire semester’s work.

Although quizzes generally help your grade point average, the exams and final exam carry a significant amount of weight toward your final average. **You must have a passing average for exams and the final exam to pass the course.** Your laboratory grade will not change a failing course grade to a passing grade.

An average less than 60 is a failing grade.

My policy is that no one will miss a grade by one point. If your actual average falls at 89, and 90 is an “A”, then your grade will be rounded up to a 90. An 88.9 will be a “B”. The same applies to the other grade ranges.

You are welcome, at any time of the semester, to stop in my office and review your course grade.

Incomplete (I) grade:

"I" grades must be requested in writing by the student. Final decisions regarding an incomplete grade are made by the instructor and are subject to review by the Department Chair and the Division Dean. Generally, the student must have successfully completed at least 2/3 of the course material to receive an "I" grade.

Incomplete grades are generally reserved for medical and family emergencies that are of significant duration or occur at a critical time during the semester, they are not a way to withdraw if you are failing the course. Please contact your instructor before the last week of class to be sure that there is sufficient time to consider your request.

An incomplete grade generally implies that a student has completed a substantial portion of the course and has shown sufficient **initiative to complete the course on his or her own**. The student will receive a copy of the standard "I" form filed with the grade. This form will detail specifically what must be done to complete the course. A student has one year to complete the required work, otherwise the grade automatically reverts to an "F."

An incomplete grade does **not** permit you to sit through the course a second time. There are liability issues regarding non-registered students in the classroom.

Late Withdrawal grade:

Your instructor has the ability to enter a late withdrawal grade (W) at the end of the semester as a final grade for the course if you request it and have not withdrawn by the normal two-thirds mark of the semester. Once final grades have been entered, you cannot retroactively request a D or F to be changed to a W.

You must request a W grade, in writing, no later than the last day of class. (Not the final exam week.)

If you have not attended class, and have not withdrawn by the 2/3 semester date, your instructor will not honor a late withdrawal request.

Please be aware that a W may affect your financial aid or other funding. You are advised to check with the funding organization to make sure there are no adverse effects to a W grade.

Final Grades:

Students do not receive a grade transcript from the college mailed to their home address at the end of the semester. Students must log on to Banner Online Services to retrieve their grade information or may check grades by calling MAX 2000 at 206-4880. For privacy and security reasons, instructors may not post grades and may **NOT** give grades over the telephone.

SOME SUGGESTIONS FOR SUCCEEDING IN CHEMISTRY

Chemistry is a complex subject. It includes a great number of abstract concepts along with mathematical manipulations of equations and data. **The material in this course is cumulative**, as you progress through the course, new material presented builds on earlier material previously covered. There is also a large vocabulary of technical terms.

As a college level course, the class moves through topics at a rapid pace with information supplementing the textbook presented in lectures. It is expected that students will read and review information in their textbooks and will consult additional sources (books, articles, and the Internet) to aid in learning and understanding course information. Because concepts and problem solving can be varied in many ways and be applied to different types of compounds and situations, **this is not a course where you can memorize course material** and solutions to specific problems.

Succeeding in chemistry requires **constant study and review** of class material since many concepts build upon previous material. **Problem solving**, actually working out problems, as listed in assignments, is essential to success in this course. This is not a course where one can just read over or cram information before a quiz or exam.

The following suggestions will be helpful in helping you to succeed in this course:

1. Always read the chapter in the textbook before it is covered in class.

Before a thorough reading, **scan the chapter(s)** to find out the kind of material covered. Look at any learning objectives or chapter outlines, the section headings, illustrations and tables, margin notes, and boxes containing relevant information and applications. Also look at the chapter summary and types of questions asked at the end of the chapter.

Read the chapter. This is technical information with its own vocabulary. You probably will not understand all the material after the first reading, it may take several readings and you may need to look up the meanings of occasional terms. *Make notations of anything you do not understand* in the textbook or on a separate sheet of paper - you will **not** remember them unless you write them down. It is in the lecture portion of the course where this material will be explained.

In addition, your instructor may use PowerPoint presentations. These presentations may be posted on the Internet. PowerPoint presentation may contain additional information not in the textbook and may only contain basic problem solving as many instructors prefer to solve problems in detail on the white board or chalk board in class.

When necessary, go back to previous chapters and review pertinent information that forms the foundation of the current material.

2. Attend the lectures.

Each topic and chapter will be explained in lecture along with illustrations, relevant applications, supplementary material, and demonstrations. Important concepts will be stressed. Often, concepts will be explained in a different way from your textbook and they may be illustrated by demonstrations. In addition, relevant material and applications, not discussed in your textbook, will be presented in the lectures. Any and all of this material may appear on exams.

Keep your notes from your chapter reading handy during the lecture to check that all your questions or uncertainties are addressed. If you are unsure of any concept, ask a question.

If you miss a class, you may miss a significant amount of course material and explanations of concepts and problem solving techniques. Although some of the material from the previous class may be briefly reviewed in the following class, new material will be covered.

3. Ask questions.

The only *stupid questions* are those that are **not** asked. If you do not understand something, chances are that there are others with the same question - **ASK IT**, no one else will.

If you are really self-conscious about asking questions in class, then ask your instructor before or after class. Also, stop in your instructor's office during his/her office hours or make an appointment to meet with your instructor.

4. Keep your work organized.

Well organized material is easier to follow and understand and it aids in your understanding of course material. Organize your notes by topic and sub-topic or rewrite them in outline form. Make notations in your notes of things you do *not* understand. This organization is helpful when reviewing for quizzes and exams.

Good organization is especially helpful with problem solving. As you work out each problem, you should show the formula used or a concept map of the solution, identify what is being asked, and list the given data and additional factors

or information needed. Substitute the proper terms into the formula and be sure to include the proper physical units. Do the arithmetic last. Use of the physical units in solving a problem provide a quick means for checking your results.

You will be asked to show clear organized set-ups for all calculations on quizzes and exams. If you are missing the formulas, the physical units, or any needed parts of the set-up for a calculation, you will be penalized by the loss of points.

5. Try all the homework questions and problems.

You must read and answer assigned questions. It is also essential that you physically work out assigned problems and calculations within a reasonably short time after they are covered in class so you will know if you have any difficulties with the material.

If you get stuck on a problem, no matter how little or how much you have done of it, *do not spend more than 10 minutes* with it, try another problem. Try the problem again, from scratch, the following morning or evening. If you are still stuck, then write a notation on your paper stating what your difficulty is (e.g. "what do I do next?"; "how do I use this item of data?", etc...). At the earliest possible opportunity, ASK your instructor or a tutor for help. **SAVE YOUR WORK** (including your rough notes) - it will give your instructor a starting point for an explanation and help you to see your error or difficulty.

You will not be able to solve problems on quizzes and exams if you wait until the night before the test to study and solve problems.

6. Study and review the course material on a regular basis.

Try to study in short sessions. You will retain more information from several 10 or 15 minute study sessions than you will from one long session. Make up some cards or summary sheets listing items that you have difficulty remembering, carry them with you, and review them when you find yourself unoccupied (such as waiting for someone, standing in a line, etc...). It is also helpful to make up mnemonics for lists of items or terms, or for concepts.

Small study groups with several of your classmates may help in understanding course material and problem solving. Generally, at least one member of the study group will understand each topic.

7. Get the addresses and telephone numbers of TWO classmates.

If you miss a class, you can get the notes from one or both classmates. Study together, you will be surprised at how much you learn from each other. Your instructor does not maintain a set of comprehensive course notes which you can copy.

One of the best ways to get help in the course is to talk to your instructor. Make an appointment with him/her and **keep it**.

Caveats:

Your instructor will make every attempt to follow the above procedures and schedules, but they may be changed in the event of extenuating circumstances.

Students submitting assignments are advised to make copies for their own protection.

If you move during the semester, please file a change of address form at any PCC campus registration office.